

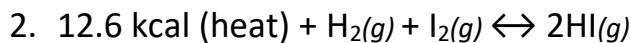
SCH4U 4-1F: Chemical Equilibrium

Applying Le Chatelier's Principle in Chemical Equilibrium

Use Le Chatelier's Principle to **determine** in which direction equilibrium will shift in each scenario. Complete the following charts by writing left, right, or none for the equilibrium shift, and decreases, increases, or remains the same for the concentrations of reactants and products. **Justify** your reasoning.



Stress	Equilibrium Shift	Reasoning
Add N_2		
Remove H_2		
Increase Temperature		
Increase Pressure		



Stress	Equilibrium Shift	Reasoning
Remove H_2		
Adding a Catalyst		
Increase Temperature		
Decrease Pressure		



Name: _____



Stress	Equilibrium Shift	Reasoning
Decrease Temperature		
Add CO_2		
Remove CO		
Increase Pressure		